***A Little Stoichiometry Fun With Magnesium…***

Name:\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

You and your partner are going to experimentally determine the stoichiometry of a chemical system. Please read the procedure carefully, wear safety goggles, and tie back any long hair.



**Materials**

* Bunsen burner
* Lighter
* Tongs
* Electronic balance
* 1 strip of magnesium ribbon

**Procedure**

* Weigh the strip of magnesium ribbon and record its mass
* Using the tongs, light the strip of magnesium ribbon on fire. \*\* Do not look directly at the light!
* Write down your observations

**Questions**

1. The mass of our magnesium ribbon was \_\_\_\_\_\_\_\_\_\_\_
2. Describe what happened when you burned the magnesium: \_\_\_\_\_\_\_\_\_\_\_\_
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3. Write out a balanced equation for the reaction that took place.
4. Find the number of moles for each product and reactant (there should be 3 in total).
5. Find the grams for each product and reactant.
6. Find the litres of oxygen that were needed for this reaction to occur.