**Stoichiometry Extra Practice**

1. Predict the products of the following reaction, then balance the reaction.

Ca + Al2(SO3)3 🡪

2. Predict the products of the following reaction, then write the balanced chemical reaction.

Sulphurous acid reacts with Iron(III)hydroxide

3. Predict the products of the reaction of nitrogen gas and hydrogen gas and write the balanced chemical equation.

4. Consider the following reactants:

K + Cl2 🡪

Complete and balance the chemical reaction, then answer the following question.

How many moles of K are needed to react with 2.45 moles of Cl2?

5. Consider the following unbalanced reaction:

KClO3  🡪 KCl + O2

How many grams of KCl will be produced from the decomposition of 2.3 moles of KClO3?

6. How many moles of carbon dioxide and how many moles of water can be made from the combustion of 10.0 grams of C3H8?

7. When water decomposes, it turns into hydrogen gas and oxygen gas. How many grams of oxygen gas will be produced from the decomposition of 10.0 grams of water?

8. How many molecules of ammonia, NH3, can be made from 10 grams of nitrogen gas, given enough hydrogen gas?